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Solution: OH OH OH 0.50 mmol mL - - - = x = nc V x 25.0 mL OH n - = 12.5 mmol HCl(aq) HCl(aq) [HCl(aq)] 12.5 mmol = = n V 0.10 mmol HCl(aq) /mL V = 120 mL 2. (a) Given: V HCl(aq) = 25.00 mL; [HCl(aq)] = 0.350 mol/L; V NaOH(aq) = 5.0 mL; [NaOH(aq)] = 0.500 mol/L; Required: [H+(aq)], pH Analysis: Before titration, amount of H+ is: H(aq) H(aq) H(aq) 0.350 mmol mL ++ += x = nc V

Section 8.7: Acid–Base Titration Tutorial 1 Practice, page 547.

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Solution: Q= [H 2 (g)][I 2 (g)] [HI(g)]2 = (0.040)(0.010) (0.14)2 Q=0.020 Statement: The value of Q is 0.020. Q is equal to K, so the reaction is at equilibrium. (b) Given: [HI(g)] = 0.20 mol/L; [H 2(g)] = 0.15 mol/L; [I 2(g)] = 0.090 mol/L; K = 0.020 Required: Q; direction of reaction to reach equilibrium Solution: Q= [H 2 (g)][I 2 (g)] [HI(g)]2 = (0.15)(0.090) (0.20)2 Q=0.34

Section 7.5: Quantitative Changes in Equilibrium Systems...

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